

**PERIODIC
TABLE**

H																	He
Li	Be											B	C	N	O	F	Ne
Na	Mg											Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

ENGAGING, HANDS-ON CHEMISTRY ACCESSIBLE TO ALL LEARNERS



For students who:

- are emerging or non-readers
- take alternate assessments
- are in special education
- short-attention span
- lack pre-requisite skills
- benefit from the use of pictures for support
- middle/high school

What's Included

-  17 days of Lesson Plans
-  51-Slide PowerPoint & MP4 Video
-  Vocabulary Activities
-  Worksheets & Graphic Organizers:
 -  Circle Maps
 -  Labeling activities
 -  Ordering elements
 -  Sorting elements
 -  Assessments
 -  Digital Activities

**Scroll
down
to see
MORE!!**

Topics Covered

- Organization of the periodic table
- Periods and groups (families)
- Metals, nonmetals, metalloids, and noble gases
- Element names and symbols
- Atomic number
- Atomic mass
- Valence (outer) electrons
- Periodic trends
- Why atoms bond
- Ionic, covalent, and polar covalent bonds



Super Simple Option

LOOKING FOR SOMETHING EASIER? CHECK OUT THIS UNIT ON
SUPER SIMPLE PERIODIC TABLE

→ **SUPER SIMPLE** ←
PERIODIC TABLE

H																			He
Li	Be											B	C	N	O	F	Ne		
Na	Mg											Al	Si	P	S	Cl	Ar		
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Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og		
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu			
		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr			

 **SPECIAL ED**

CLICK HERE



Lesson Plans

The **lesson plans** contain:

- ✓ Overall tips for teaching students with significant needs
- ✓ A quick look at what you will do each day
- ✓ Detailed instructions on how that day's lesson should run

Quick Look

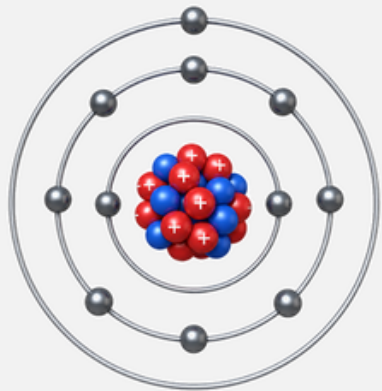
Day	Activity	Day	Activity
1	- Book - Vocab cards activity - Circle map	10	- Book - Vocab cards activity - Valence electrons activity
2	- Book - Vocab cards activity - Circle map	11	- Book - Vocab cards activity - Valence electrons activity
3	- Book - Vocab cards activity - Sorting activity	12	- Book - Who am I activity - Compare types of
4	- Book - Vocab cards activity - Sorting activity	13	- Book - Who am I activity - Compare types of
5	- Book - Vocab cards activity - Sorting activity	14	- Book - Who am I activity - Compare types of
6	- Book - Vocab cards activity - Labeling activity	15	- Book - Who am I activity - Sorting bonds
7	- Book - Vocab cards activity - Labeling activity	16	- Book - Bingo - Close works
8	- Book - Vocab cards activity - Ordering activity	17	- Book - Bingo - Assessment
9	- Book - Vocab cards activity - Ordering activity		

Day 12

Activity	Notes	Materials
Read or listen to a recording of the PowerPoint (10 minutes)	- Read through the story, asking lots of questions - Continue to make connections between PowerPoint and vocabulary board	- PowerPoint - Vocabulary board
Who am I activity (10 minutes)	- Give each student a set of Who am I cards - Various options/ways to play, but the goal is to have students hold up the card you are talking about. Sometimes there can be more than one correct answer Things you can do: - Hold up a vocabulary card - Show a page from the book - Point to a symbol on the vocabulary board - Verbally describe something about that card - See teacher instructions with cards for more ideas.	Who am I cards (2 sets)
Valence electron activity review (5 minutes)	Review the activity completed yesterday	Activity completed yesterday
Comparing types of bonds (15 minutes)	- Students will first color in a diagram of a bond (ionic, covalent, or polar) according to the directions. - Using what they learned, students will explain/sequence the steps that occur when that bond forms	- Worksheet - Scissors - Glue - Crayons or colored pencils
Sharing (10 minutes)	Each student shares their activity with the group using the communication method of their choice	- Completed puzzle - Communication devices

PowerPoint

The mass number tells about how much the atom weighs. It depends mostly on the number of protons and neutrons.



● = Proton (11)
● = Neutron (12)
● = Electron (11)

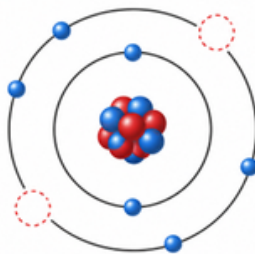
SODIUM
Atomic Number: 11
Electron Configuration: 2-8-1

This is a sodium atom. It has 11 protons and 12 neutrons. It has a **mass number** of 23.

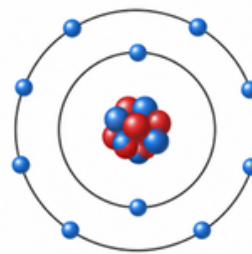
$$\begin{array}{r} 11 \\ + 12 \\ \hline 23 \end{array}$$

Most atoms do not like having an outer shell that is only partly full. They become more stable when their outer shell is full, or when it has 8 electrons.

Not Stable
Outer shell is only partly full.



Stable
Outer shell is full.



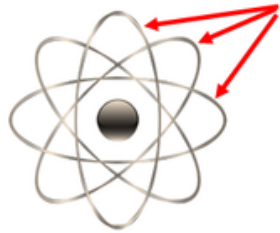
51-slide **PowerPoint** using photos and simple text.

- ✓ PowerPoint
- ✓ Google Slides
- ✓ mp4 video file

Vocabulary

orbital

The path electrons travel around the nucleus.



ductile

Ability of metals to be pulled into thin wires.



ionic bond

Bond between two atoms in which one atom gives up an electron and the other takes it.



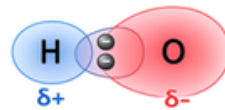
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covalent bond

Bond between two atoms where electrons are shared.

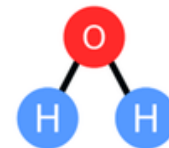
polar bond

A covalent bond in which the electrons are not shared equally.



chemical bond

Invisible glue that holds atoms together. They can be covalent or ionic.



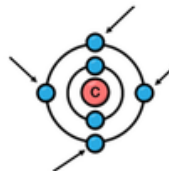
Noble gases

Elements in the far right group of the periodic table are very stable. They have a full outer electron shell.



valence electrons

Electrons in the outermost shell.

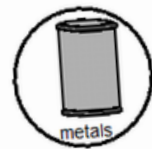


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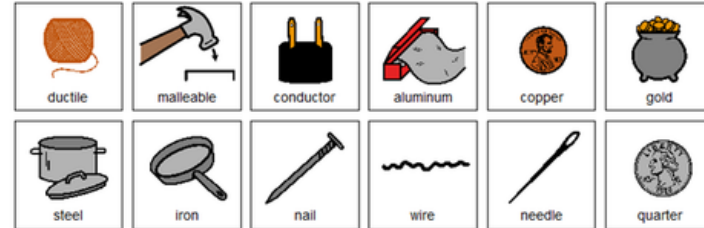
 This unit comes with a vocabulary board and vocabulary cards.

- Improve **participation**.
- Increase **engagement** in group discussions.
- Daily group activities.

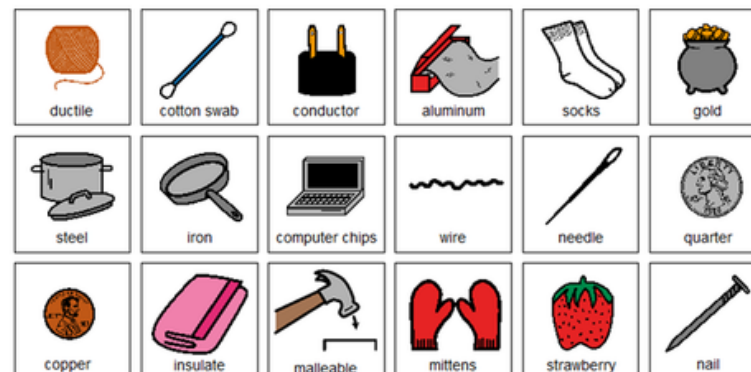
3 Circle maps



Place the following in the circle map on previous page.



Place the following in the circle map on previous page
ONLY IF correct.



- ✓ Periodic table
- ✓ Metals
- ✓ Nonmetals

- **ERRORLESS VERSION**
- **WRONG ANSWERS MIXED IN**

Labeling Activities

Color in a **period** red. Color in a **group** blue.

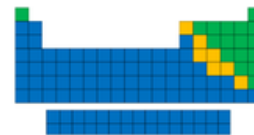
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				57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu
				89 Ac	90 Th	91 Pa	92 U	93 Np										

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- ✓ Periods and groups
- ✓ Metals, metalloids, nonmetals
- ✓ Noble gases
- ✓ Hydrogen and Helium

Color in:

- Metals blue
- Metalloids orange
- Nonmetals green

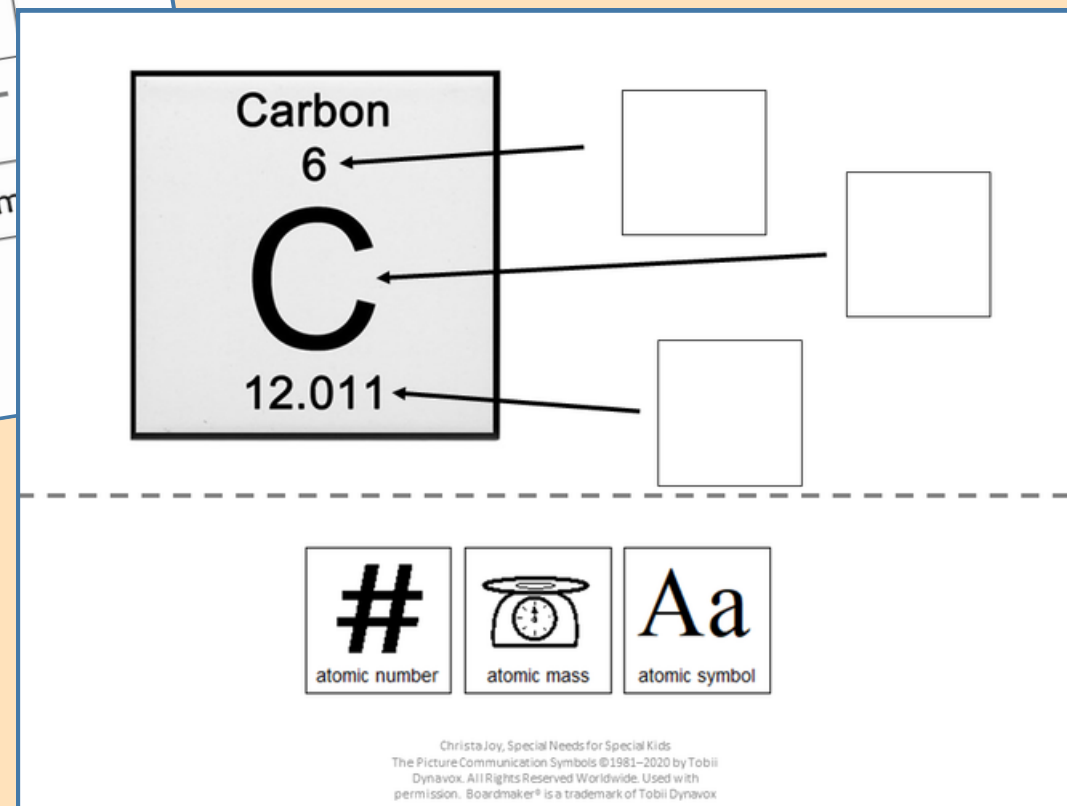
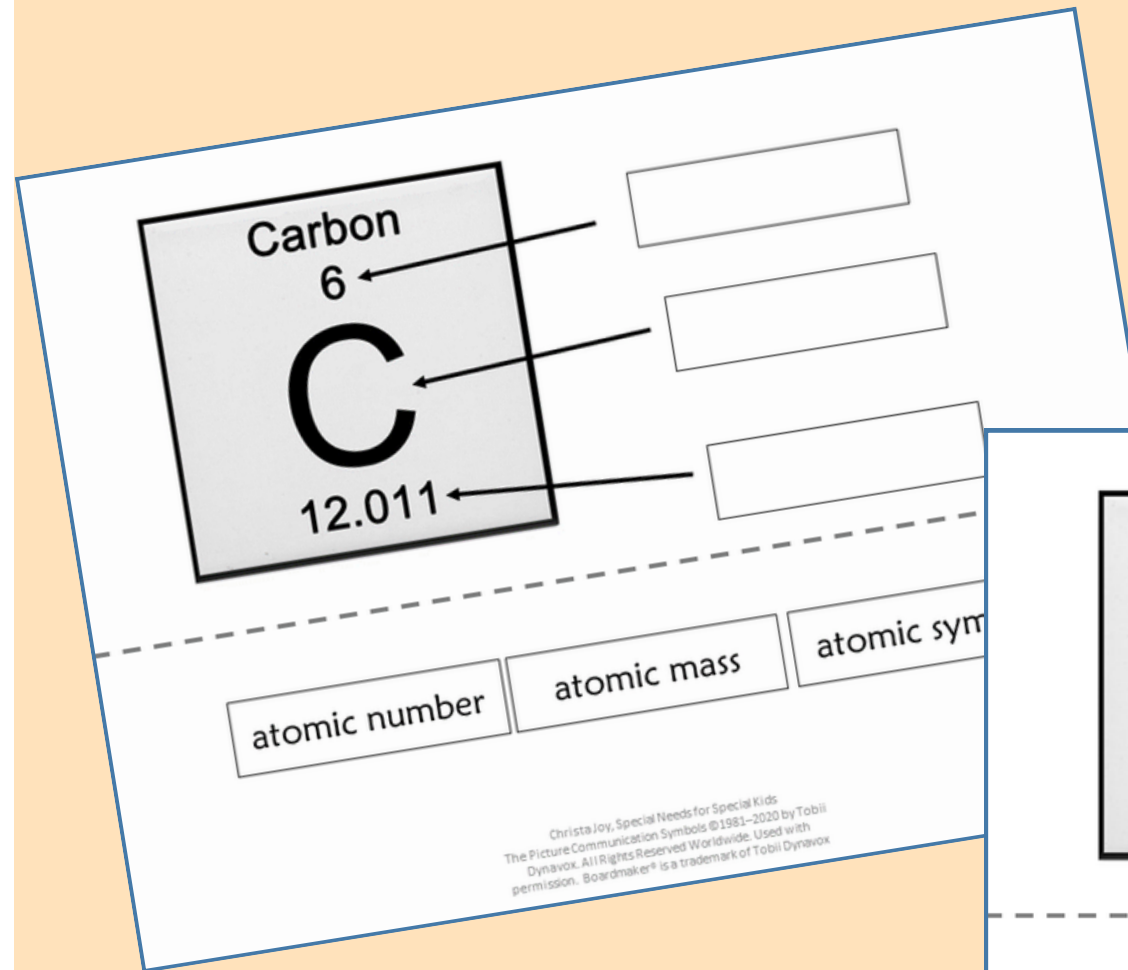


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				89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

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Reading Atomic Symbols

USING WORDS OR PICTURES



Ordering Elements

1. Cut out the elements on the bottom of the page and order them from least to greatest atomic number.
2. Color those same three elements on the periodic table.

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8	9	5
O	F	B

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PRACTICE READING ATOMIC SYMBOLS

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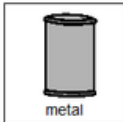
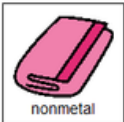
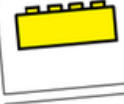
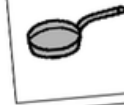
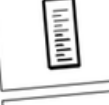
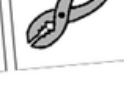
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38	51	46
Sr	Sb	Pb

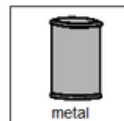
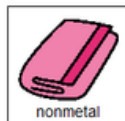
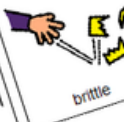
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Sorting Activities

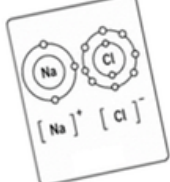
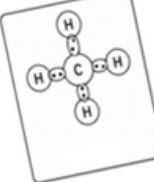
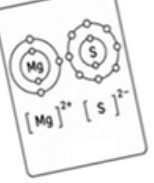
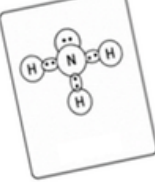
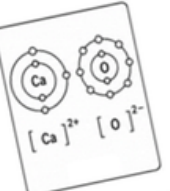
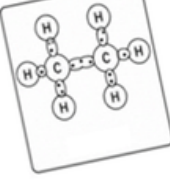
IN COLOR AND BLACK & WHITE

 metal  nonmetal				
				
				
				
				

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 metal  nonmetal				
				
				

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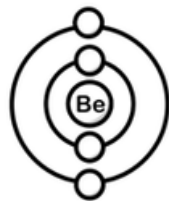
 Ionic bond	 Covalent bond
 	 
 	 

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Trends in the table

1. Count the number of valence electrons in each atomic model.
2. Color those same three elements on the periodic table.

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Be



Mg



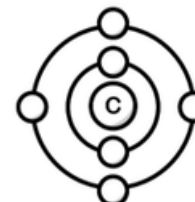
Ca



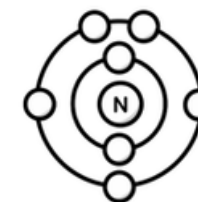
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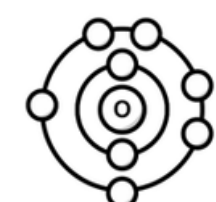
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5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
6	55 Cs	56 Ba		72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
7	87 Fr	88 Ra		104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Uut	114 Uuq	115 Uup	116 Uuh	117 Uus	118 Uuo
Lanthanides			57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	
Actinides			89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	



C



N



O



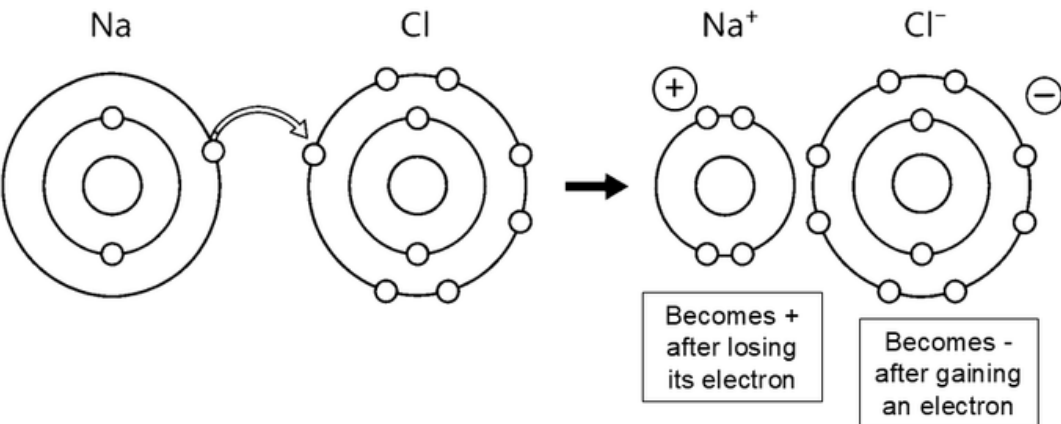
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Forming Bonds

- 1. IONIC BONDS
- 2. COVALENT BONDS
- 3. POLAR BONDS

Forming Ionic Bonds

- 1. Color the nucleus pink.
- 2. Color the electrons blue.
- 3. Trace the path of the electron moving red.

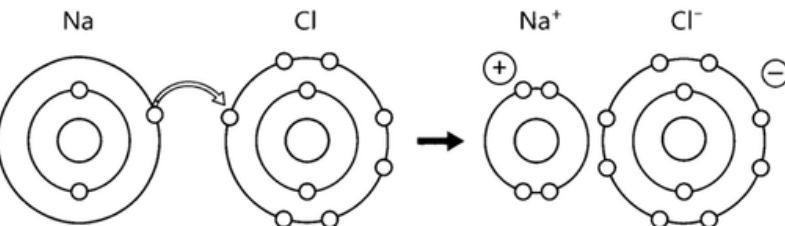


Images for ionic bond.

 gives electron	 positive	 opposite charges attract
 negative	 more stable	 takes electron

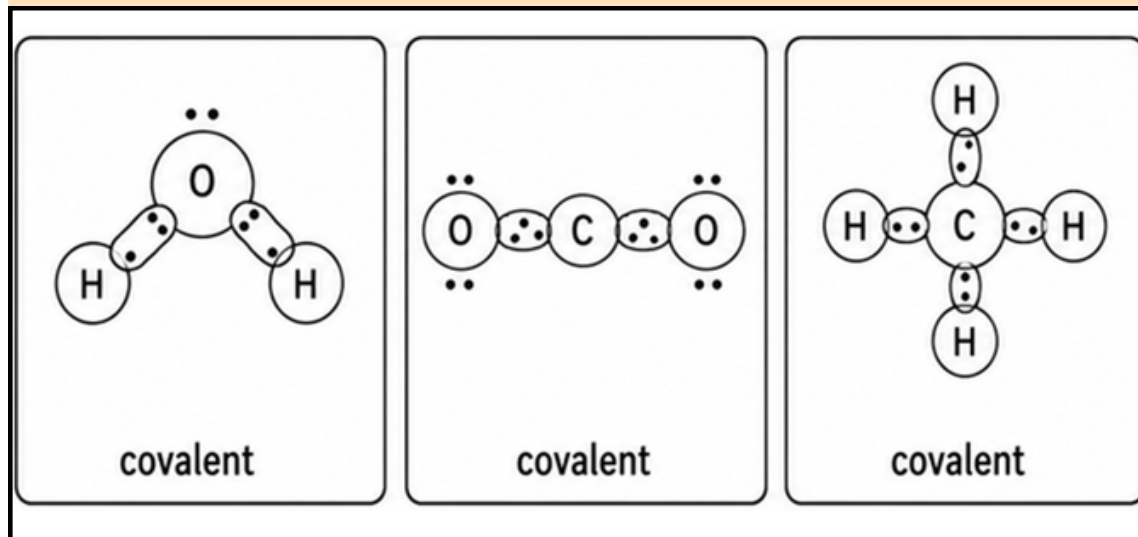
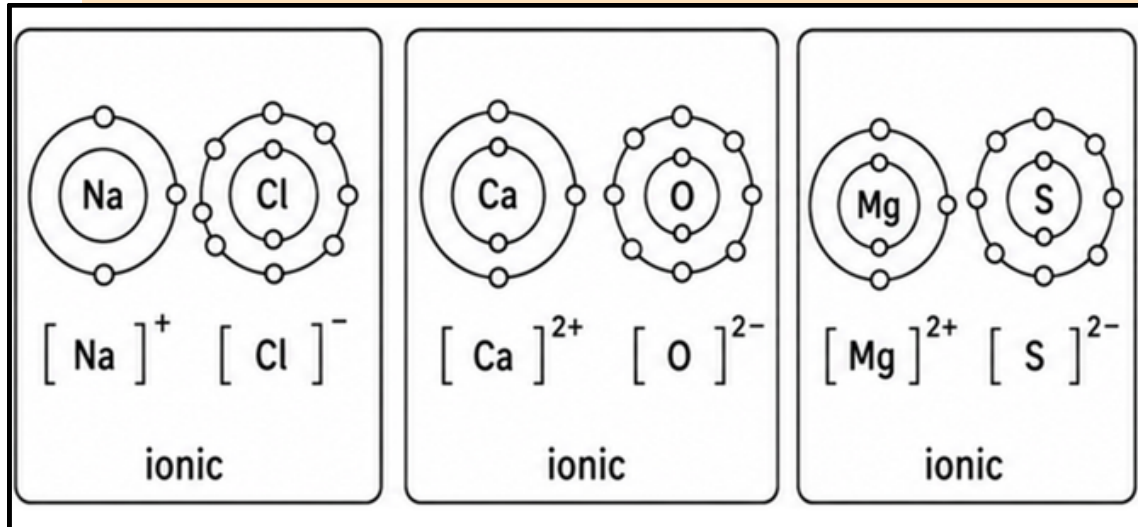
Forming Ionic Bonds

Trace the steps in forming an ionic bond.

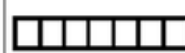
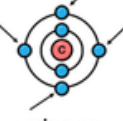
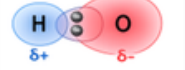
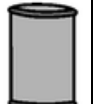


What happens first?	What happens next?	Why do they stay together?	What is the result?

Group Activities



Periodic Table

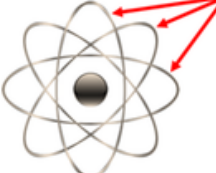
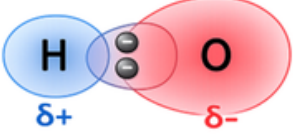
 ductile	 period	 electronic configuration	$\#$  atomic number
 covalent bond	 group	 mass number	 valence electron
 polar bond	 orbit	 noble gas	 human
 metalloids	 nonmetals	 metals	

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1. VOCABULARY REVIEW

2. WHO AM I?

3. BINGO

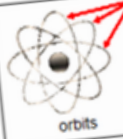
 orbit	 ionic bond
 covalent bond	 polar bond

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Assessment

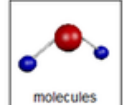
The Periodic Table

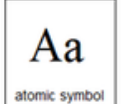
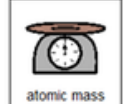
- The periodic table organizes all the .
- Ionic bonds electrons.
- Covalent bonds electrons.
- The atomic number tells you how many are in the nucleus.
- Each period has elements with the same number of .

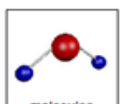
 share
  orbits
  protons
  give
  elements

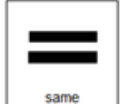
Version 1

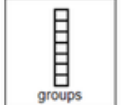
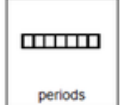
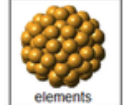
- The periodic table is an organization of .
- What is the abbreviation for each element?
- The atomic mass tells you how many are in the nucleus.
- As you go across a period (left to right), what happens to the atomic mass?
- What are the vertical columns in the periodic table called?

 solids
  elements
  molecules

 atomic symbol
  atomic number
  atomic mass

 electrons
  molecules
  protons

 decreases
  same
  increases

 groups
  periods
  elements

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Quiz

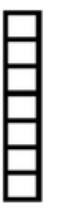


Assessment

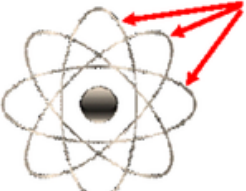
Version 2

Print onto cardstock or mount on index cards. Cut pictures apart and show student answer choices for each question.

Q 5

 groups
  periods
  elements

Q 6

 protons
  orbits
  atomic mass

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Digital Activities



- ✓ Provide extra practice
- ✓ Great independent work centers
- ✓ Include a differentiated set of slides
- ✓ Interactive

Day 4

Move red circles over elements in one group. Move blue circles over elements in one period.

Day 7 differentiated

LABELING ELEMENTS WITH PICTURES
Move the correct label on the right to the correct target on the left.

Carbon
6
C
12.011

atomic mass

Aa
atomic symbol

atomic number

Christa Joly, Special Needs for Special Kids
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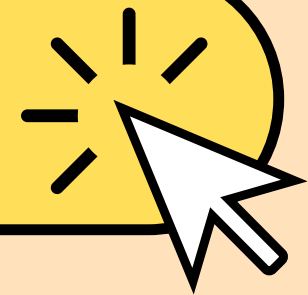
What teachers are saying

This was a great tool to teach grade-level science content at a level that was accessible to my students.

I love this activity for my special needs students. This is a great resource for my classroom.

**SAVE MONEY AND GET THIS AS PART OF
THE CHEMISTRY BUNDLE.**

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Christa Joy, M.Ed., DVM